



Mark Scheme (Results)

January 2023

Pearson Edexcel International GCSE
In Chemistry (4CH1)
Paper 2CR

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a)	7		1
(b)	any value ≥ 4 but < 7		1
(c)	0	ALLOW Group 8	1
(d)	3		1

(Total for Question 1 = 4 marks)

Question number	Answer	Notes	Marks
2 (a)	temperature	ALLOW volume/amount of solvent IGNORE concentration IGNORE stirring	1
(b) (i)	F → G → B → D → H → C → A → E M1 G in correct place M2 H C A in correct places		2
(ii)	M1 mass of dry salt (= $78.1 - 60.5$ =) 17.6 (g) M2 solubility = $(17.6 \times 2) = 35.2$ (g per 100 g of water)	ALLOW ECF from M1 correct answer with no working scores 2	2

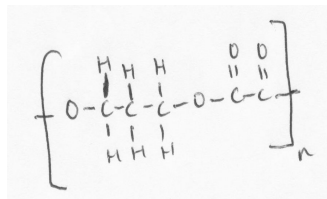
(Total for Question 2 = 5 marks)

Question number	Answer	Notes	Marks
3 (a)	M1 calculating heat energy in kJ/g $= (3.28 \times 10^5 \div 10\,000)$ OR 32.8 M2 calculating heat energy in kJ/mol $= (M1 \times 12)$ OR 393.6 M3 = 394 OR 3.94×10^2 kJ/mol OR M1 calculating amount of carbon $= 10\,000 \div 12$ OR 833.3 M2 calculating heat energy in kJ/mol $= (3.28 \times 10^5 \div M1)$ OR 393.6 M3 = 394 OR 3.94×10^2 kJ/mol	IGNORE sign M3 subsumes M2 IGNORE sign M3 subsumes M2 correct answer with no working scores 3	3
(b) (i)	M1 $\frac{600}{20\,000} \times 100$ OR $\frac{0.6}{20} \times 100$ M2 = 3 (%)	ALLOW ECF from incorrect conversion of units in M1 correct answer with no working scores 2	2
(ii)	M1 mol of sulfur = $600 \div 32$ OR 18.75 M2 vol SO₂ = $M1 \times 24$ OR 450 (dm³) M3 = 450 000 OR 4.5×10^5 (cm³)	ALLOW ECF from incorrect A_r in M1 ALLOW ECF M2 (dm³) × 1000 M3 subsumes M2 correct answer with no working scores 3	3
(iii)	acid rain	IGNORE any effects	1

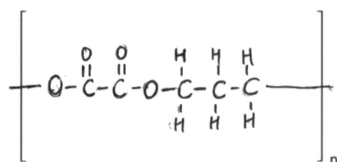
(Total for Question 3 = 9 marks)

Question number	Answer	Notes	Marks
4 (a) (i)	The correct answer is D a structural formula A is incorrect because $\text{CH}_3\text{CH}_2\text{OH}$ is not a displayed formula B is incorrect because $\text{CH}_3\text{CH}_2\text{OH}$ is not an empirical formula C is incorrect because $\text{CH}_3\text{CH}_2\text{OH}$ is not a molecular formula		1
(ii)	Any two from the following: M1 same general formula M2 each member differs from the next by CH_2 M3 same functional group M4 similar chemical properties/reactions M5 trend/gradation in physical properties OWTTE	ACCEPT named physical property eg boiling point	2
(iii)	alcohol(s)		1
(iv)	propan-1-ol	ALLOW 1-propanol ALLOW propanol	1
(b) (i)	explanation including M1 condensation (polymerisation) M2 because water also produced/eliminated/removed/lost	ALLOW small molecule also produced/eliminated/removed/lost	2

(ii)



OR



M1 correct ester linkage shown

M2 rest of repeat unit fully correct

M2 DEP M1

2

(Total for Question 4 = 9 marks)

Question number	Answer	Notes	Marks
5 (a)	2.8.5		1
(b)	M1 formula of phosphide ion is P^{3-} M2 charges on three calcium/Ca^{2+} ions balance / cancel out charges on two phosphide/P^{3-} ions OWTTE	ALLOW P^{3-} ALLOW charge on phosphide ion is 3- / -3 ALLOW reference to three calcium atoms each lost two electrons / (total of) six electrons which were gained by two phosphorus atoms (to attain full outer shells) OWTTE Any reference to sharing of electrons/covalent bonding scores 0	2
(c) (i)	$\text{Ca}_3(\text{PO}_4)_2 + 8 \text{ C} \rightarrow \text{Ca}_3\text{P}_2 + 8 \text{ CO}$		1
(ii)	explanation including M1 (carbon acts as a) reducing agent M2 (because) calcium phosphate/$\text{Ca}_3(\text{PO}_4)_2$ loses oxygen	ALLOW (because) carbon gains oxygen / is oxidised ALLOW carbon removes oxygen from calcium phosphate/$\text{Ca}_3(\text{PO}_4)_2$ IGNORE references to reactivity series / displacement reactions M2 DEP M1 correct or missing	2
(d)	$\text{Ca}_3\text{P}_2 + 6\text{H}_2\text{O} \rightarrow 3\text{Ca}(\text{OH})_2 + 2\text{PH}_3$ M1 all formulae correct M2 correct balancing	M2 DEP M1 ALLOW multiples and fractions	2

(e)	explanation including M1 giant (ionic) structure M2 strong forces of attraction between oppositely charged ions M3 which require a lot of (heat/thermal) energy to break/overcome	ACCEPT giant (ionic) lattice ACCEPT between Ca^{2+} and P^{3-} ACCEPT between positive and negative ions ALLOW strong ionic bonds IGNORE more energy 0 marks if any mention of covalent bonds, intermolecular forces or molecules	3
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(Total for Question 5 = 11 marks)

Question number	Answer	Notes	Marks
6 (a) (i)	explanation including M1 carbon above iron in reactivity series M2 so carbon can remove oxygen from iron(III) oxide	ALLOW carbon is more reactive than iron ACCEPT reverse arguments ALLOW carbon can reduce iron(III)oxide ALLOW carbon can displace iron (from iron(III) oxide)	2
(ii)	electricity not been discovered OWTTE	IGNORE electrolysis not discovered	1
(b) (i)	explanation including M1 ions M2 can move	If refs to electrons moving then scores 0 M2 DEP M1	2
(ii)	The correct answer is C Na^+ A is incorrect because H^+ ions not present in molten sodium chloride B is incorrect because Cl^- ion is not a cation D is incorrect because OH^- ion not present in molten sodium chloride		1
(iii)	explanation including M1 water/moisture reacts with sodium M2 to produce hydrogen (which ignites/burns/reacts in oxygen/air causing the small explosions)		2
(iv)	M1 (anode) $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^{(-)}$ M2 (cathode) $(2)\text{Na}^+ + (2)\text{e}^{(-)} \rightarrow (2)\text{Na}$	ACCEPT $2\text{Cl}^- - 2\text{e}^{(-)} \rightarrow \text{Cl}_2$ both half-equations correct but order reversed scores 1	2

Question number	Answer	Notes	Marks
6 (c)	explanation including M1 outer electron in potassium (atom) is further away (from nucleus) M2 outer electron in potassium (atom) less strongly attracted by nucleus OWTTE M3 so (outer) electron more easily lost OWTTE	ALLOW reverse arguments ALLOW potassium (atom) has more (electron) shells than lithium (atom) ALLOW potassium (atom) larger than lithium (atom) ALLOW potassium has greater atomic radius	3

(Total for Question 6 = 13 marks)

Question number	Answer	Notes	Marks
7 (a)	to (more) easily/clearly see the colour change (at end point) OWTTE		1
(b)	M1 litmus M2 blue OR M1 methyl orange M2 yellow OR M1 phenolphthalein M2 pink	M2 DEP M1 ALLOW purple ALLOW orange ALLOW any other suitable indicator and correct final colour	2
(c)	Description including any five from the following M1 rinse/wash the conical flask (with distilled/deionised water) M2 (repeat titration/experiment) adding sodium hydroxide (from burette) slowly/dropwise (near end point) M3 swirling flask M4 record initial and final volume burette reading (at end point) /record volume sodium hydroxide added (at end point) M5 repeat until obtain concordant results M6 find mean/average (of concordant results)	REJECT if rinsed/washed using solution IGNORE names of indicators and any colour changes ALLOW shaking/stirring ALLOW reference to subtraction of initial and final readings ALLOW results within 0.2 cm ³ (or less)	5
(d) (i)	moles of NaOH = $\frac{(0.350 \times 18.80)}{(1000)} = 0.00658$	ACCEPT 6.58×10^{-3}	1
(ii)	amount of HNO ₃ = 0.00658	ALLOW ECF from (i)	1
(iii)	conc. of HNO ₃ = $\frac{0.00658 \times 1000}{25.0} = 0.263(2)$	ALLOW ECF from (ii) If not divided by 1000 in (i) do not penalise if not multiplied by 1000 in (iii) ALLOW 2, 3 or 4 sig figs throughout Penalise use of 1 sig fig once only	1

(Total for Question 7 = 11 marks)

Question number	Answer	Notes	Marks
8 (a)	diagram including the following M1 horizontal lines showing energy levels labelled $\text{N}_2 + \text{O}_2$ and 2NO M2 level of 2NO /products above level of $\text{N}_2 + \text{O}_2$ /reactants M3 ΔH correctly shown between reactants and products and labelled M4 activation energy correctly shown and labelled	ALLOW vertical line with/without arrowheads If single arrowhead must point from level of reactants to level of products Must be from level of reactants to top of "hump" ALLOW vertical line with/without arrowheads If single arrowhead must point from level of reactants to top of "hump" IGNORE any label on a horizontal axis If diagram for exothermic reaction drawn can score M1 M3 M4	4
(b) (i)	$\{944 + (3 \times 436) =\}$ 2252	IGNORE any sign	1
(ii)	$\{6 \times 391 =\}$ 2346	IGNORE any sign	1
(iii)	M1 difference between (i) and (ii) expected value = 94 M2 – 94	IGNORE any sign ALLOW ECF from (i) and (ii) If (ii) > (i) sign should be – If (i) > (ii) sign should be +	2

(Total for Question 8 = 8 marks)
(Total for Paper = 70 marks)

