



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/32

October/November 2021

1 hour 15 minutes

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

- 1 (a) A list of formulae is shown.

CaCO_3
 CaO
 Cl_2
 CH_4
 $\text{C}_2\text{H}_5\text{OH}$
 C_2H_6
 CuSO_4
 H_2
 H_2O
 MgO
 NaCl
 SO_2

Answer the following questions using these formulae.
Each formula may be used once, more than once or not at all.

State which formula represents:

- (i) a compound that is the main constituent of natural gas

..... [1]

- (ii) an element that is used in water treatment

..... [1]

- (iii) an element that bleaches damp litmus paper

..... [1]

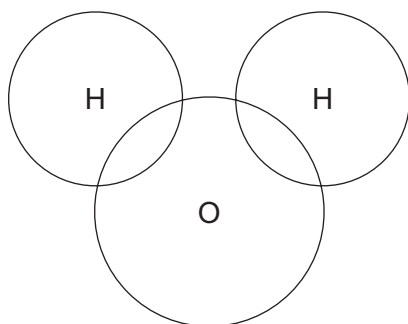
- (iv) a compound that contains an ion with a single negative charge

..... [1]

- (v) a hydrocarbon that is formed by the decomposition of vegetation.

..... [1]

(b) Complete the dot-and-cross diagram to show the electron arrangement in a molecule of water.



[2]

(c) State whether calcium oxide is a basic oxide or an acidic oxide.
Give a reason for your answer.

.....
..... [1]

[Total: 8]

- 2 The table shows the masses of some of the ions in 1000 cm^3 of water taken from a lake.

name of ion	formula of ion	mass of ion in 1000 cm^3 of lake water / mg
calcium	Ca^{2+}	0.41
chloride	Cl^-	4.40
magnesium	Mg^{2+}	0.39
	NO_3^-	0.03
potassium	K^+	0.30
silicate	SiO_3^{2-}	0.02
	Na^+	2.90
sulfate	SO_4^{2-}	2.80

- (a) Answer these questions using only the information in the table.

- (i) State which of the negative ions has the lowest concentration.

..... [1]

- (ii) Name the compound containing Na^+ and NO_3^- ions.

..... [1]

- (iii) Calculate the mass of chloride ions in 250 cm^3 of lake water.

mass = mg [1]

- (b) Describe a test for sulfate ions.

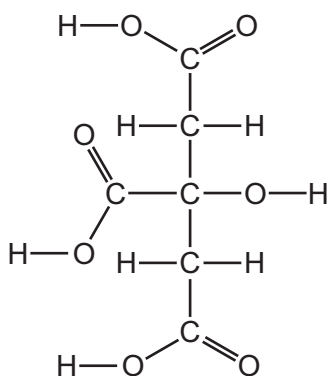
test

observations

[2]

(c) Citric acid is also present in the lake water.

The structure of citric acid is shown.



(i) Deduce the number of carboxylic acid groups in one molecule of citric acid.

..... [1]

(ii) The formula of citric acid is $C_6H_8O_7$.

Complete the table to calculate the relative molecular mass of citric acid.

type of atom	number of atoms	relative atomic mass	
carbon	6	12	$6 \times 12 = 72$
hydrogen		1	
oxygen		16	

relative molecular mass = [2]

[Total: 8]

3 Iron is extracted from iron ore in a blast furnace.

(a) Name an ore of iron.

..... [1]

(b) (i) Complete the chemical equation for the reduction of iron(III) oxide in the blast furnace.



(ii) State the meaning of the term *reduction*.

..... [1]

(c) Calcium carbonate (limestone) is added to the blast furnace.
The calcium carbonate undergoes thermal decomposition.

State the meaning of the term *thermal decomposition*.

.....
..... [2]

(d) Iron can be made into stainless steel.

(i) Give **one** use of stainless steel.

..... [1]

(ii) Describe **one** advantage of stainless steel compared with pure iron.

.....
..... [1]

(e) The symbol for an isotope of iron is shown.



Deduce the number of electrons, neutrons and protons in one atom of this isotope of iron.

number of electrons

number of neutrons

number of protons

[3]

(f) Iron is a good conductor of heat and electricity.

Give two **other** physical properties of iron that are characteristic of **all** metals.

1

2 [2]

(g) Iron rusts.

Name the **two** substances needed for iron to rust.

1

2 [2]

[Total: 15]

4 The table shows some properties of the Group I elements.

element	melting point / °C	density in g/cm ³	observations during reaction with water
lithium	181	0.53	
sodium	98		rapid bubbling no flame
potassium		0.86	rapid bubbling lilac flame
rubidium	39	1.53	very rapid bubbling red flame
caesium	29	1.88	explodes
francium	27		

(a) (i) Complete the table by predicting:

- the melting point of potassium
- the density of francium.

[2]

(ii) Describe the observations when lithium reacts with water.

.....
 [1]

(b) (i) Deduce the electronic structure of sodium.
 Use the Periodic Table to help you.

..... [1]

(ii) Explain why a potassium ion has a single positive charge.

.....
 [1]

- (c) Sodium reacts with water to produce aqueous sodium hydroxide and a gas which 'pops' with a lighted splint.

(i) Complete the chemical equation for this reaction.



(ii) Choose **one** value from the list that best describes the pH of aqueous sodium hydroxide.

Draw a circle around the correct answer.

pH 1 pH 4 pH 7 pH 14 [1]

[Total: 8]

5 The table shows the structures of some organic compounds.

compound	structure of compound	homologous series
G	$ \begin{array}{ccccc} & \text{H} & & \text{H} & & \text{H} \\ & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{H} \\ & & & & & \\ & \text{H} & & \text{H} & & \text{H} \end{array} $	alkane
H	$ \begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{O} - \text{H} \\ & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & \end{array} $	
J	$ \begin{array}{ccccc} & \text{H} & & & & \text{H} \\ & & & & & \\ \text{H} & - \text{C} & - & \text{C} & = & \text{C} \\ & & & & & \\ & \text{H} & & \text{H} & & \text{H} \end{array} $	

- (a) Complete the table by naming the homologous series.
The first one has been done for you.

[2]

- (b) Draw the structure of a compound containing two carbon atoms which belongs to the same homologous series as compound **H**.
Show all of the atoms and all of the bonds.

[1]

- (c) Describe the colour change when an excess of compound **J** is added to aqueous bromine.

from to [2]

- (d) (i) Compound **J** can be obtained by cracking petroleum fractions.

State the conditions needed for cracking.

.....
 [2]

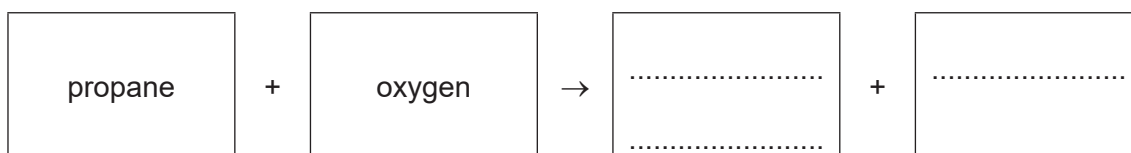
- (ii) Complete this sentence about cracking using a word from the list.

bitumen hydrogen oxygen petroleum

The chemicals manufactured by cracking include alkanes, alkenes and [1]

- (e) Compound **G** is propane.

Complete the word equation for the complete combustion of propane.



[2]

- (f) Compound **J** can form polymers.

- (i) State the meaning of the term *polymer*.

.....
 [2]

- (ii) Nylon is also a polymer.

Give **one** use of nylon.

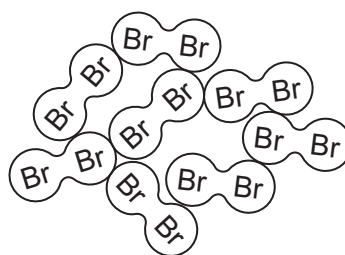
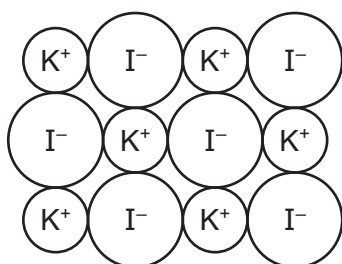
..... [1]

- (iii) Describe **one** pollution problem caused by non-biodegradable plastics.

.....
 [1]

[Total: 14]

- 6 The diagrams show part of the structures of potassium iodide and bromine at room temperature and pressure.



- (a) Describe the physical properties of these substances in terms of:

- volatility

potassium iodide

bromine

- solubility in water

potassium iodide

bromine

- electrical conductivity when molten (liquid).

potassium iodide

bromine

[5]

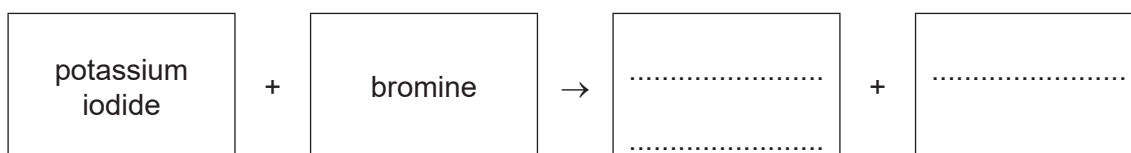
- (b) Molten potassium iodide is electrolysed using carbon (graphite) electrodes.

- (i) Name the substance produced at the positive electrode.

..... [1]

- (ii) Aqueous potassium iodide reacts with aqueous bromine.

Complete the word equation for this reaction.



[2]

- (iii) Explain in terms of the reactivity of the halogens why aqueous potassium chloride does **not** react with aqueous bromine.

.....

..... [1]

- (c) Name the change of state when liquid bromine changes to solid bromine.

..... [1]

[Total: 10]

7 This question is about sulfur and compounds of sulfur.

(a) Use the kinetic particle theory to describe the differences between sulfur gas and solid sulfur in terms of:

- the arrangement of the particles

.....

.....

.....

- the separation of the particles.

.....

.....

.....

[4]

(b) Give the major use of sulfur in industry.

..... [1]

(c) Sulfur dioxide is a pollutant in the air that contributes to acid rain.

(i) State **one** adverse effect of sulfur dioxide on health.

..... [1]

(ii) Name one **other** oxide that contributes to acid rain.

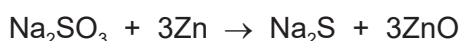
..... [1]

(iii) Sulfur dioxide reacts with water to produce sulfurous acid.
The reaction is reversible.

Draw the symbol for a reversible reaction in the box.



(d) The equation for the reaction of sodium sulfite with zinc is shown.



Explain how this equation shows that zinc is oxidised.

.....

..... [1]

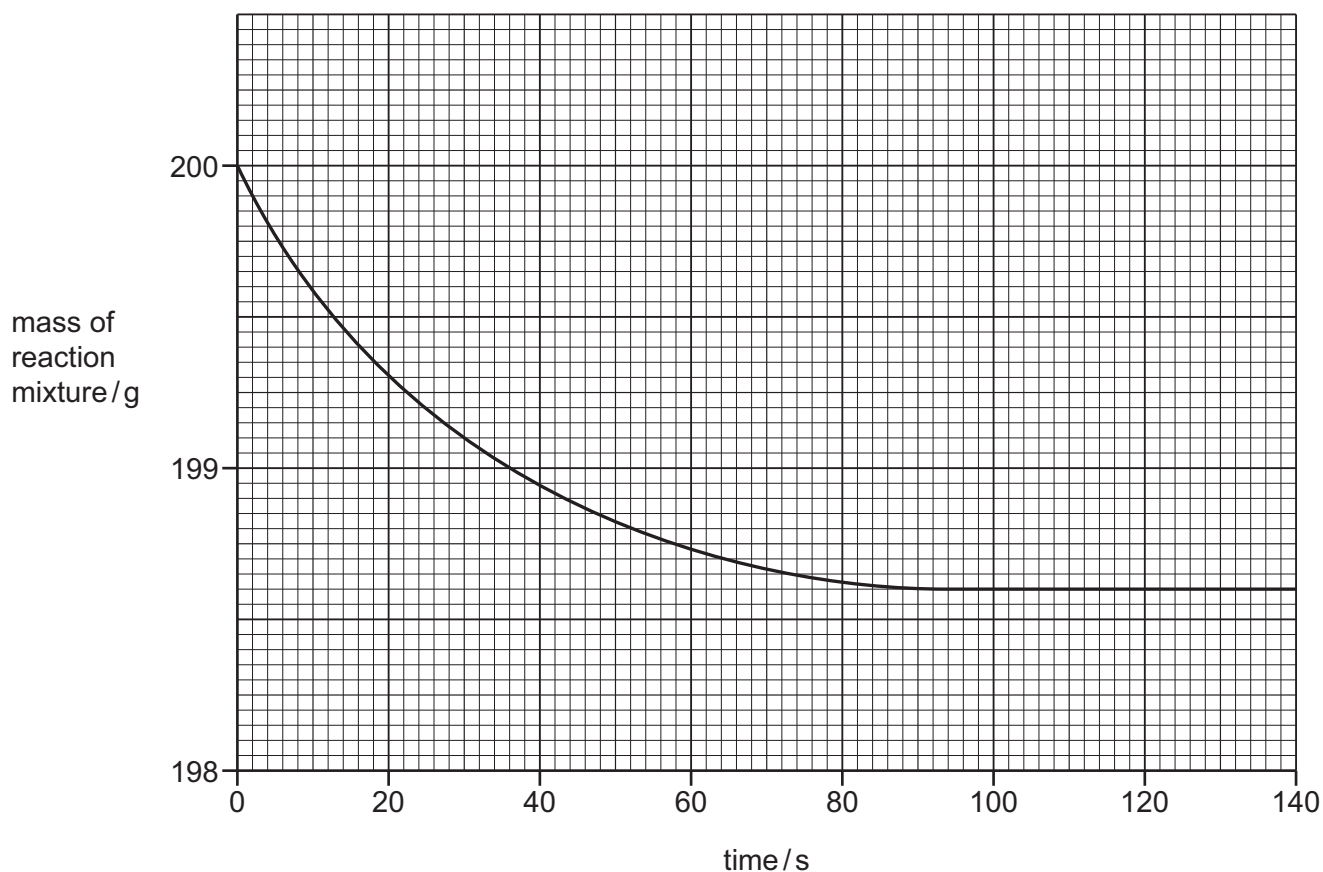
[Total: 9]

- 8 A student investigated the reaction of small pieces of calcium carbonate with dilute hydrochloric acid. The hydrochloric acid was in excess.



The rate of reaction is found by measuring the decrease in the mass of the reaction mixture with time.

The results are shown on the graph.



- (a) Deduce the time taken from the beginning of the experiment for the mass of the reaction mixture to decrease by 1.0 g.

time = s [1]

- (b) The experiment was repeated using dilute hydrochloric acid of a higher concentration.

All other conditions stayed the same.

Draw a line **on the grid** to show how the mass of the reaction mixture changes with time using acid of a higher concentration. [2]

- (c) Describe the effect each of the following has on the rate of reaction of calcium carbonate with hydrochloric acid.

All other conditions stay the same.

- The reaction is carried out at a higher temperature.

.....

- The reaction is carried out using large pieces of calcium carbonate instead of small pieces of calcium carbonate.

.....

[2]

- (d) When 0.44 g of calcium carbonate is used, 100 cm³ of carbon dioxide gas is formed.

Calculate the mass of calcium carbonate needed to produce 25 cm³ of carbon dioxide gas.

mass of calcium carbonate = g [1]

- (e) The table compares the reaction of four metals with dilute hydrochloric acid.

metal	observations
iron	bubbles produced slowly
magnesium	bubbles produced very rapidly
nickel	bubbles produced very slowly
silver	no bubbles produced

Put the four metals in order of their reactivity.

Put the least reactive metal first.

least reactive $\longrightarrow$ most reactive

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[2]

[Total: 8]

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57	La	lanthanum	139
58	Ce	cerium	140
59	Pr	praseodymium	141
60	Nd	neodymium	144
61	Pm	promethium	—
62	Sm	samarium	150
63	Eu	euporium	152
64	Gd	gadolinium	157
65	Tb	terbium	159
66	Dy	dysprosium	163
67	Ho	holmium	165
68	Er	erbium	167
69	Tm	thulium	169
70	Yb	ytterbium	173
71	Lu	lutetium	175
72	Hf	hafnium	178
73	Ta	tantalum	181
74	W	wolfram	184
75	Re	rhodium	186
76	Os	osmium	190
77	Ir	iridium	192
78	Pt	platinum	195
79	Au	gold	197
80	Hg	mercury	201
81	Tl	thallium	204
82	Pb	lead	207
83	Bi	bismuth	209
84	Po	polonium	—
85	At	astatine	—
86	Rn	radon	—
87	Fr	francium	—
88	Ra	radium	—
89	Ac	actinium	—
90	Th	thorium	232
91	Pa	protactinium	231
92	U	uranium	238
93	Np	neptunium	—
94	Pu	plutonium	—
95	Am	americium	—
96	Cm	curium	—
97	Bk	berkelium	—
98	Cf	californium	—
99	Es	einsteinium	—
100	Fm	fermium	—
101	Md	mendelevium	—
102	No	nobelium	—
103	Lr	lawrencium	—
104	Rf	rutherfordium	—
105	Db	bohrium	—
106	Sg	seaborgium	—
107	Bh	bohrium	—
108	Hs	hassium	—
109	Mt	meitnerium	—
110	Ds	darmstadtium	—
111	Rg	roentgenium	—
112	Cn	copernicium	—
113	Nh	nihonium	—
114	Fl	flerovium	—
115	Mc	moscovium	—
116	Lv	livermorium	—
117	Ts	tennessine	—
118	Og	oganeson	—

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).